

## Reactions

## 4. Types of Reaction

## CONCEPT 2

## TEST YOURSELF

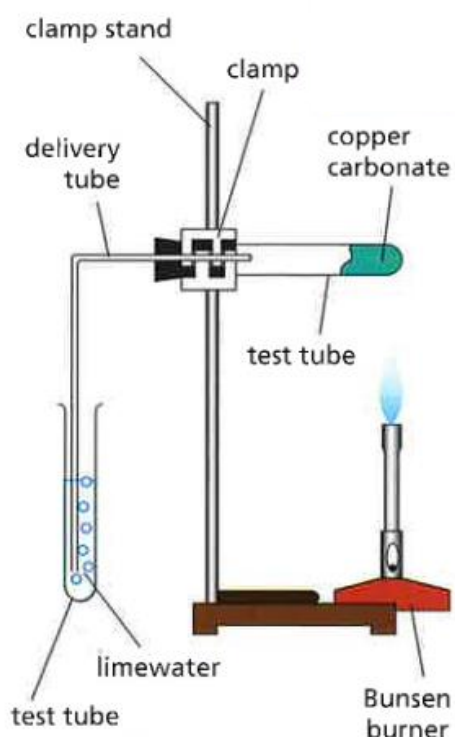
## THERMAL DECOMPOSITION

## KNOW

- Q1 What are the reactants and products of the thermal decomposition of a metal carbonate?
- Q2 Why is thermal decomposition a good name for this type of reaction?

## APPLY

- Q3 Which elements does calcium carbonate contain?
- Q4 In the experiment in the diagram, what safety precautions should be taken?
- Q5 Describe the changes you would expect to see in this experiment?
- Q6 Before it was heated, the test tube and copper carbonate had a mass of 50g. Explain why, after the experiment, the mass was less than 50g.



## EXTEND

- Q7 Explain why calcium carbonate is thought to be one of the most useful carbonates.
- Q8 Write a word equation for the reaction between calcium oxide and water.
- Q9 Suggest which carbonate would be easier to decompose, magnesium carbonate or calcium carbonate? Explain your answer.